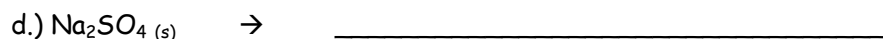


## Solutions Part 1

Name - \_\_\_\_\_

1.) Write an equation for the dissociation of each of the following in water.



2.) Which of the above solutions are electrical conductors?

3.) If 1.00 L of a 1.00 M solution of  $\text{AgNO}_3$  was mixed, then;

a.)  $[\text{Ag}^+] =$

b.)  $[\text{NO}_3^-] =$

4.) If 500. mL of a 1.00 M solution of  $\text{BaCl}_2$  was mixed, then;

a.) How many moles of  $\text{Ba}^{+2}$  are present?

b.)  $[\text{Ba}^{+2}] =$

c.) How many moles of  $\text{Cl}^-$  are present?

d.)  $[\text{Cl}^-] =$

5.) If 500. mL of 1.00 M  $\text{NaCl}$  was added to the solution in question 4, then;

a.)  $[\text{Ba}^{+2}] =$

b.)  $[\text{Cl}^-] =$