Solutions Part 1

Name - _____

1.) Write an equation for the dissociation of each of the following in water.

a.) $BaCl_{2\ (s)}$ → _____

b.) $AgNO_{3\ (s)}$ → _____

c.) $Mg(OH)_{2\ (s)}$ → _____

d.) $Na_2SO_{4\ (s)}$ → _____

e.) $NH_4NO_{3\ (s)}$ → _____

f.) $(NH_4)_3PO_{4\ (s)}$ → _____

2.) Which of the above solutions are electrical conductors?

3.) If $1.00\ L$ of a $1.00\ M$ solution of $AgNO_3$ was mixed, then;

a.) $[Ag^+]$ =

b.) $[NO_3^-]$ =

4.) If $500.\ mL$ of a $1.00\ M$ solution of $BaCl_2$ was mixed, then;

a.) How many moles of Ba^{+2} are present?

b.) $[Ba^{+2}]$ =

c.) How many moles of Cl^- are present?

d.) $[Cl^-]$ =

5.) If $500.\ mL$ of $1.00\ M$ NaCl was added to the solution in question 4, then;

a.) $[Ba^{+2}]$ =

b.) $[Cl^-]$ =