

Solutions - More Practice

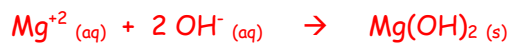
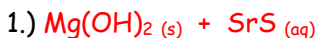
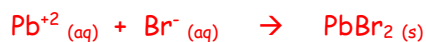
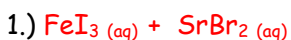
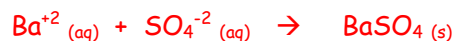
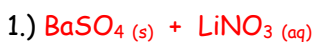
Name - _____

1.) For each of the following combinations of equal volumes of 0.2 M aqueous solutions,

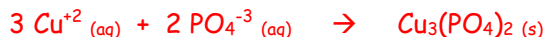
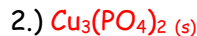
1.) Identify the possible products by formula.

2.) State which (if any) product has a low solubility (will precipitate).

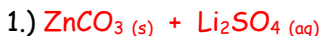
3.) If there IS a precipitate write the formula equation and net ionic equation for the reaction.

a.) $\text{MgS} + \text{Sr}(\text{OH})_2$ b.) $\text{CuBr}_2 + \text{Pb}(\text{NO}_3)_2$ c.) $\text{FeBr}_3 + \text{SrI}_2$ d.) $\text{Ba}(\text{NO}_3)_2 + \text{Li}_2\text{SO}_4$ 

e.) $K_3PO_4 + CuCl_2$



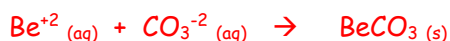
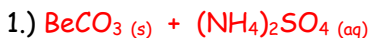
f.) zinc (II) sulphate and lithium carbonate



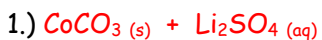
g.) iron (III) nitrate and magnesium sulphide



h.) beryllium sulphate and ammonium carbonate



i.) cobalt (II) sulphate and lithium carbonate



j.) magnesium sulphate and strontium hydroxide

