

SCIENCE 10 HONOURS NAMES AND FORMULAS

1. Write the formulas of the following compounds.

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|-----------------------|-------|-----------------------|-------|
| a. potassium bromide | _____ | f. sodium iodide | _____ |
| b. magnesium chloride | _____ | g. calcium bromide | _____ |
| c. calcium oxide | _____ | h. potassium oxide | _____ |
| d. aluminum iodide | _____ | i. beryllium chloride | _____ |
| e. barium fluoride | _____ | j. aluminum oxide | _____ |

2. Write the names of the following compounds.

- | | | | |
|----------------------|-------|----------------------|-------|
| a. NaCl | _____ | f. CaF ₂ | _____ |
| b. BaO | _____ | g. K ₂ S | _____ |
| c. Li ₂ S | _____ | h. BeBr ₂ | _____ |

3. Write the formulas or names for the following compounds containing polyatomic ions.

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|--------------------------------------------------|-------|--------------------------------------|-------|
| a. calcium acetate | _____ | f. aluminum hydroxide | _____ |
| b. sodium phosphate | _____ | g. ammonium phosphate | _____ |
| c. lithium dichromate | _____ | h. Al(NO ₃) ₃ | _____ |
| d. H ₂ CO ₃ | _____ | i. NH ₄ Cl | _____ |
| e. K ₂ Cr ₂ O ₇ | _____ | j. calcium hydroxide | _____ |

4. The metals in the following compounds form ions with different charges. Write the names and formulas for the compounds.

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|-------------------------|-------|--------------------------------------|-------|
| a. tin (IV) chloride | _____ | h. SnO | _____ |
| b. tin (II) chloride | _____ | i. SnBr ₄ | _____ |
| c. tin (IV) oxide | _____ | j. FeS | _____ |
| d. tin (II) oxide | _____ | k. FeCl ₃ | _____ |
| e. iron (II) phosphate | _____ | l. Cu ₂ CO ₃ | _____ |
| f. iron (II) fluoride | _____ | m. Sn(NO ₃) ₂ | _____ |
| g. copper (II) sulphate | _____ | n. NiC ₂ O ₄ | _____ |

5. Write the names for the following covalent compounds.

a. SiC _____

e. SO₂ _____

b. CS₂ _____

f. SO₃ _____

c. SF₆ _____

g. N₂O₅ _____

d. OF₂ _____

h. NO _____

6. Write the formulas for the following covalent compounds.

a. chlorine dioxide _____

d. nitrogen triiodide _____

b. dichlorine monoxide _____

e. diphosphorus tetraoxide _____

c. iodine tribromide _____

f. dihydrogen monoxide _____