6.1 Types of Chemical Reactions

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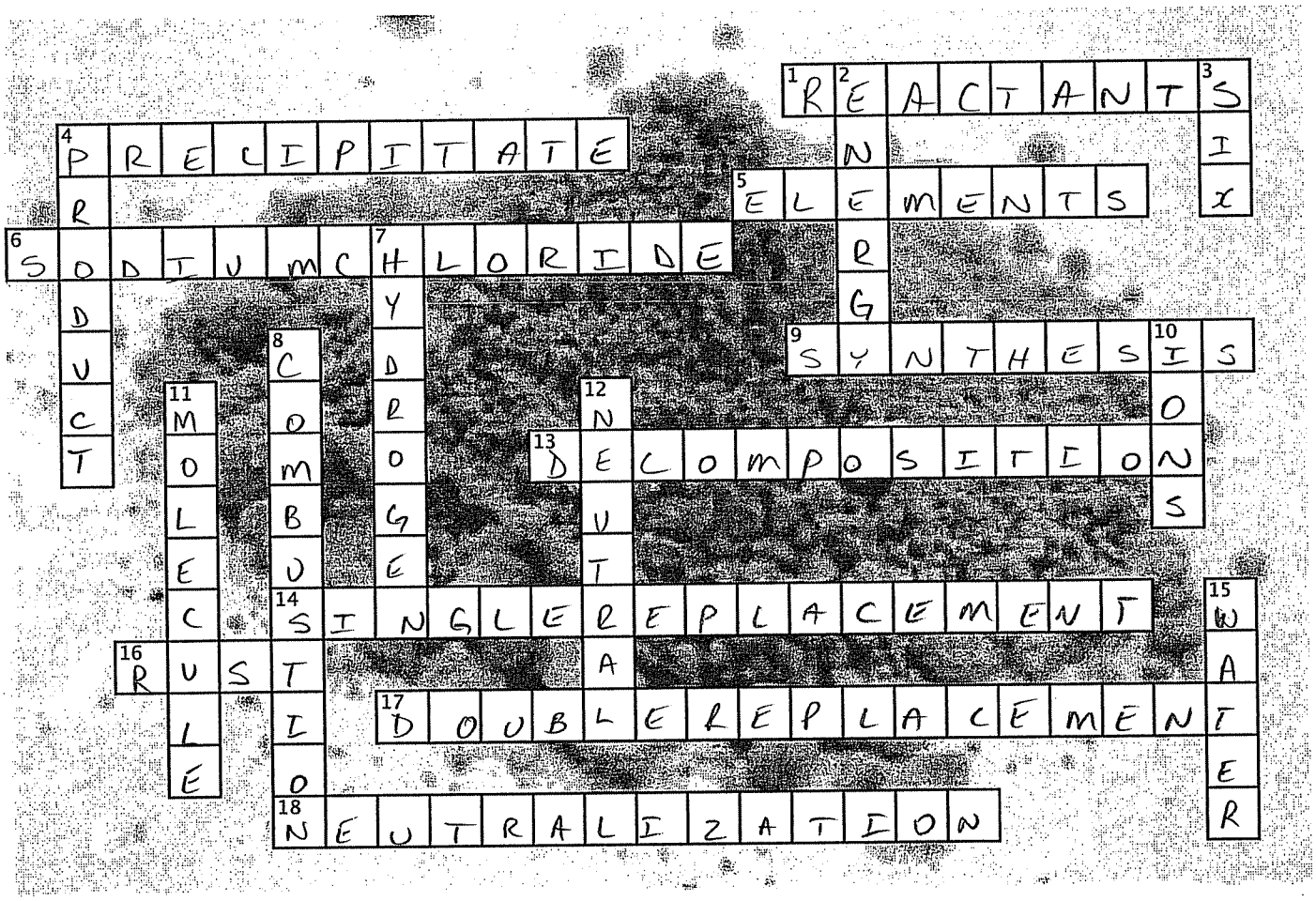

Across

1. You can identify each type of chemical reaction by examining the _____.
4. An insoluble solid that forms from a solution.
5. When two or more reactants (A and B) combine to produce a single product (AB), for example, the letters A and B represent
6. When sodium hydroxide solution is mixed with iron(III) chloride, a precipitate occurs involving the iron(III) ion. This is a double replacement reaction producing iron hydroxide and _____
9. Two or more reactants (A and B) combine to produce a single product (AB).
13. This type of reaction is the reverse of a synthesis reaction
14. A reactive element (a metal or a nonmetal) and a compound react to produce another element and another compound.
16. When iron reacts with oxygen, _____ is produced.
17. A _____ _____ reaction usually involves two ionic solutions that react to produce two new ionic compounds.
18. In a _____ reaction, an acid and a base react to form a salt and water.

Down

2. All known chemical reactions require _____ to break the chemical bonds in the reactants.
3. Chemists have identified _____ common types of reactions.
4. For ionic compounds, you can use the ion charges to predict the _____
7. Zinc metal reacts with hydrochloric acid to produce zinc chloride and _____ gas.
8. The rapid reaction of a compound or element with oxygen to form an oxide.
10. When synthesis reactions occur between a metal and non-metal, electrons are transferred from the metal to the non-metal, producing _____.
11. To make table salt in a synthesis reaction, two atoms of sodium metal and one _____ of chlorine gas react to form sodium chloride, NaCl.
12. During decomposition of an ionic compound, electrons transfer back to the atoms of the metal and each element becomes electrically _____
15. When a hydrocarbon and oxygen combust, the products are two oxides, _____ and carbon dioxide.