KEY

Naming Chemicals

Name - _____

1.) Write either the chemical formula or chemical name for the binary (two atom types) ionic compounds:

a.) strontium sulphide _____SrS_____

b.) $CaCl_2$ _____calcium chloride_____

c.) lithium bromide _____LiBr_____

d.) K_2O _____potassium oxide_____

e.) beryllium oxide _____BeO_____

f.) Li_2O _____lithium oxide_____

g.) potassium phosphide _____K_3P_____

h.) Ca_3N_2 _____calcium nitride_____

2.) Write the name or formula for these MULTIVALENT metal ionic compounds:

a.) molybdenum (III) sulphide _____Mo_2S_3_____

b.) $PbCl_2$ _____lead (II) chloride_____

c.) rhodium (IV) bromide _____$RhBr_4$_____

d.) Cu_2O _____copper (I) oxide_____

e.) mercury (II) oxide _____HgO_____

f.) Tl_2Se _____thallium (I) selenide_____

g.) bismuth (V) phosphide _____Bi_3P_5_____

h.) PbF_4 _____lead (IV) fluoride_____

3.) Write the chemical formula or name for the following compounds containing POLYATOMIC IONS:

a.) sodium acetate _____$NaCH_3COO$_____

b.) $PbCrO_4$ _____lead (II) chromate_____

c.) barium acetate _____$Ba(CH_3COO)_2$_____

d.) $CaSO_4$ _____calcium sulphate_____

e.) mercury (II) chloride _____$HgCl_2$_____

f.) Li_3PO_4 _____lithium phosphate_____

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g.) bismuth (V) phosphate _____$Bi_3(PO_4)_5$_____

h.) $(NH_4)_2HPO_4$ __ammonium monohydrogenphosphate__

4.) Classify each of the following as binary, multivalent or polyatomic (combinations can exist). Then the chemical formula or name for each.

a.) potassium acetate _____polyatomic - KCH_3COO_____

b.) PbF_2 _binary and multivalent – lead (II) fluoride_

c.) calcium acetate _____polyatomic – $Ca(CH_3COO)_2$_____

d.) Li_2O _____binary – lithium oxide_____

e.) copper (II) chlorite _polyatomic and multivalent – $Cu(ClO_2)_2$_

f.) Na_3PO_4 _____polyatomic – sodium phosphate_____

g.) lithium phosphide _____binary – Li_3P_____

h.) $(NH_4)_2SO_4$ ____polyatomic – ammonium sulphate_____