

Chemical Formulas

Rule #1 : When looking at a chemical formula, the subscript numbers below and to the left of a symbol tell us how many atoms of that particular element are present in that molecule.

ex. - H_2O : Two Hydrogen atoms and one Oxygen atom per molecule

Diagram :

Rule #2 : When brackets appear in a chemical formula it tells us that the molecule is made up of a group of elements that appear more than once. The number below and to the left tells us how many times that group is found in that molecule.

ex. - $Al(OH)_3$: One Aluminum atom, three Oxygen atoms and three Hydrogen atoms per molecule

Diagram :

Give the meaning (explain what atoms would be found in such a molecule) for each of the following chemical formulas :

a) PbS : _____

b) KNO_3 : _____

c) $NaOH$: _____

d) H_2O : _____

e) $C_{10}H_8$: _____

f) H_2O_2 : _____

g) $CuSO_4$: _____

h) Fe_2O_3 : _____

i) C_2H_5OH : _____

j) $Ca(OH)_2$: _____

k) $Al_2(SO_4)_3$: _____

l) $Cu(OH)_2$: _____

m) $Zn_3(PO_4)_2$: _____

n) $Fe_2(SO_4)_3$: _____

o) $Cu(ClO_3)_2$: _____
