

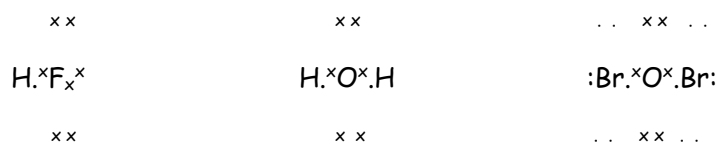
Chemical Equations - Review

- 1.) Metal atoms become ions by losing an electron or electrons so they have more protons than electrons (a positive charge).
- 2.) Non-metals become ions by gaining an electron or electrons causing them to have more electrons than protons (a negative charge).
- 3a.) Multivalent means to have more than one option of how many electrons will be lost or gained when bonding.
- b.) One metal that is multivalent could be iron.
- c.) One metal that is not multivalent is sodium.
- 4.) The symbol and name of each of the following are:
- a.) Magnesium - Mg b.) Bromine - Br c.) Gold - Au d.) Lithium - Li e.) Helium - He

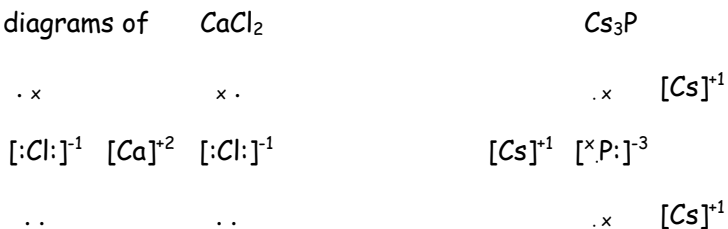
5a.) Lewis diagram of H₂ and F₂



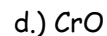
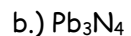
b.) Lewis diagrams of HF, H₂O, and OBr₂



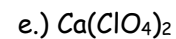
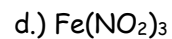
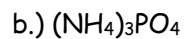
c.) Lewis diagrams of



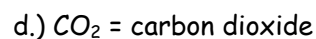
6.) Write the names of the following.



7.) Write the formula for each of the following.



8.) Write the name of each chemical.



9.) Balance the following equations.

