

Chemical Equations - Review

- 1.) Explain how metal atoms become ions.

- 2.) Explain how non-metal atoms become ions.

- 3.) Define "multivalent" and then give examples of both a multivalent metal and one metal that is not multivalent.

- 4.) Name and give the symbol for each of the following.
 - a.) The element in period 3 and group 2
 - b.) The halogen in period 4
 - c.) The element in period 6 and group 11
 - d.) The alkali metal in period 2
 - e.) The noble gas in period 1
- 5.) Draw the Lewis diagram for each of the following.
 - a.) The molecules H_2 F_2

 - b.) The compounds HF H_2O OBr_2

 - c.) The compounds CaCl_2 CS_3P

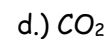
6.) Write the name for each of the following multivalent metal compounds.

- a.) Gold(III) fluoride
- b.) Lead(IV) nitride
- c.) Copper(I) iodide
- d.) Nickel(III) sulphide
- e.) Chromium(II) oxide

7.) Write the formula for each of the following.

- a.) Sodium carbonate
- b.) Ammonium phosphate
- c.) Ammonium nitrate
- d.) Iron(III) nitrite
- e.) Calcium perchlorate

8.) Write the name of each chemical.



9.) Balance the following equations.

