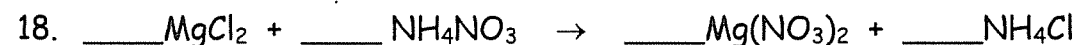
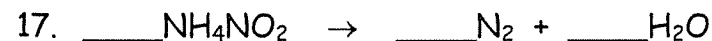
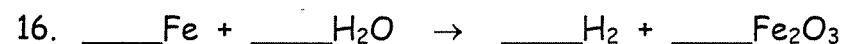
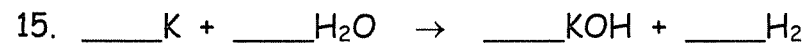
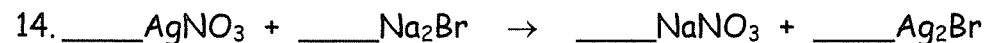
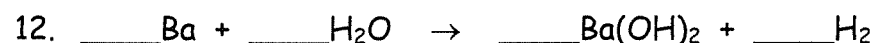
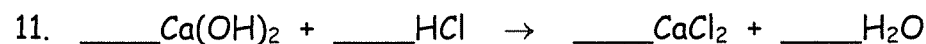
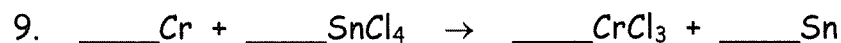
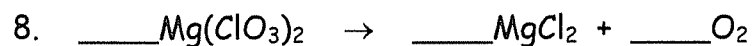
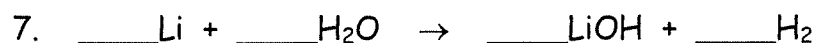
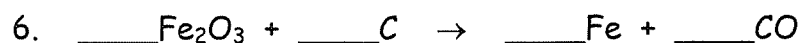
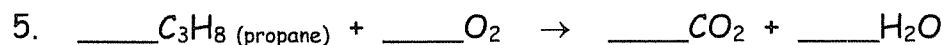
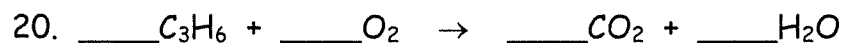
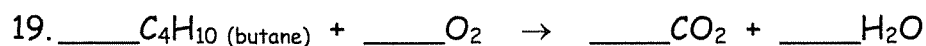


Balancing and Chemical Reactions

Balance the following equations by placing the appropriate coefficients in front of the formulas:



(A Little Tougher?!)



42. _____ $\text{Al}_2(\text{SO}_4)_3(\text{aq})$ + _____ $\text{Ca}(\text{OH})_2(\text{aq}) \longrightarrow$ _____ $\text{Al}(\text{OH})_3(\text{s})$ + _____ $\text{CaSO}_4(\text{s})$
43. _____ $\text{K}_2\text{CO}_3(\text{s})$ + _____ $\text{H}_3\text{PO}_4(\text{aq}) \longrightarrow$ _____ $\text{K}_3\text{PO}_4(\text{aq})$ + _____ $\text{H}_2\text{O}(\text{l})$ + _____ $\text{CO}_2(\text{g})$
44. _____ $\text{K}(\text{s})$ + _____ $\text{F}_2(\text{g}) \longrightarrow$ _____ $\text{KF}(\text{s})$
45. _____ $\text{Sr}(\text{s})$ + _____ $\text{O}_2(\text{g}) \longrightarrow$ _____ $\text{SrO}(\text{s})$
46. _____ $\text{NI}_3(\text{s}) \longrightarrow$ _____ $\text{N}_2(\text{g})$ + _____ $\text{I}_2(\text{g})$
47. _____ $\text{Ca}(\text{OH})_2(\text{s}) \longrightarrow$ _____ $\text{CaO}(\text{s})$ + _____ $\text{H}_2\text{O}(\text{g})$
48. _____ $\text{NH}_4\text{NO}_2(\text{s}) \longrightarrow$ _____ $\text{N}_2(\text{g})$ + _____ $\text{H}_2\text{O}(\text{l})$
49. _____ $\text{MoS}_2(\text{s})$ + _____ $\text{O}_2(\text{g}) \longrightarrow$ _____ $\text{MoO}_3(\text{s})$ + _____ $\text{SO}_2(\text{g})$
50. _____ $\text{Mg}(\text{OH})_2(\text{aq})$ + _____ $\text{H}_3\text{PO}_4(\text{aq}) \longrightarrow$ _____ $\text{Mg}_3(\text{PO}_4)_2(\text{aq})$ + _____ $\text{H}_2\text{O}(\text{l})$
51. _____ $\text{Pb}(\text{NO}_3)_2(\text{aq})$ + _____ $\text{K}_2\text{CrO}_4(\text{aq}) \longrightarrow$ _____ $\text{PbCrO}_4(\text{s})$ + _____ $\text{KNO}_3(\text{aq})$
52. _____ $\text{CaCO}_3(\text{s})$ + _____ $\text{HC}_2\text{H}_3\text{O}_2(\text{aq}) \longrightarrow$ _____ $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2(\text{aq})$ + _____ $\text{H}_2\text{O}(\text{l})$ + _____ $\text{CO}_2(\text{g})$
53. _____ $\text{NH}_4\text{Cl}(\text{aq})$ + _____ $\text{Pb}(\text{NO}_3)_2(\text{aq}) \longrightarrow$ _____ $\text{PbCl}_2(\text{s})$ + _____ $\text{NH}_4\text{NO}_3(\text{g})$
54. _____ $\text{C}_6\text{H}_{12}\text{O}_6(\text{aq})$ + _____ $\text{O}_2(\text{aq}) \longrightarrow$ _____ $\text{CO}_2(\text{aq})$ + _____ $\text{H}_2\text{O}(\text{l})$
55. _____ $\text{NaHCO}_3(\text{s}) \longrightarrow$ _____ $\text{H}_2\text{O}(\text{g})$ + _____ $\text{CO}_2(\text{g})$ + _____ $\text{Na}_2\text{CO}_3(\text{s})$
56. _____ $\text{Na}_2\text{CO}_3(\text{s})$ + _____ $\text{HCl}(\text{aq}) \longrightarrow$ _____ $\text{NaCl}(\text{aq})$ + _____ $\text{H}_2\text{O}(\text{l})$ + _____ $\text{CO}_2(\text{g})$
57. _____ $\text{Cl}_2(\text{aq})$ + _____ $\text{KI}(\text{aq}) \longrightarrow$ _____ $\text{KCl}(\text{aq})$ + _____ $\text{I}_2(\text{aq})$
58. _____ $\text{ZnO}(\text{aq})$ + _____ $\text{HCl}(\text{aq}) \longrightarrow$ _____ $\text{ZnCl}_2(\text{s})$ + _____ $\text{H}_2\text{O}(\text{l})$
59. _____ $\text{Mg}(\text{s})$ + _____ $\text{N}_2(\text{g}) \longrightarrow$ _____ $\text{Mg}_3\text{N}_2(\text{s})$
60. _____ $\text{C}_6\text{H}_6(\text{l})$ + _____ $\text{O}_2(\text{g}) \longrightarrow$ _____ $\text{CO}_2(\text{g})$ + _____ $\text{H}_2\text{O}(\text{g})$
61. _____ $\text{Na}(\text{s})$ + _____ $\text{H}_2\text{O}(\text{l}) \longrightarrow$ _____ $\text{NaOH}(\text{aq})$ + _____ $\text{H}_2(\text{g})$
62. _____ $\text{Al}(\text{s})$ + _____ $\text{O}_2(\text{g}) \longrightarrow$ _____ $\text{Al}_2\text{O}_3(\text{s})$
63. _____ $\text{AgNO}_3(\text{aq})$ + _____ $\text{CuCl}_2(\text{aq}) \longrightarrow$ _____ $\text{AgCl}(\text{s})$ + _____ $\text{Cu}(\text{NO}_3)_2(\text{aq})$
64. _____ $(\text{NH}_4)_2\text{SO}_4(\text{aq})$ + _____ $\text{NaOH}(\text{aq}) \longrightarrow$ _____ $\text{NH}_3(\text{g})$ + _____ $\text{H}_2\text{O}(\text{l})$ + _____ $\text{Na}_2\text{SO}_4(\text{aq})$

