

Analysis of the Double Replacement Reaction Between Na_2CO_3 and CaCl_2

Name - _____

Partner - _____

Part 1 - Day 1

Purpose - to practice techniques in conducting an experiment while completing calculations designed to practice basic stoichiometry in calculating CaCO_3 produced and yield.

Materials -

Procedure -

- 1.) Measure the mass of approximately 2.00 g of Na_2CO_3 by first placing a clean 250 mL beaker on the electronic balance, taring the balance (that means to "zero" the balance), and carefully adding about 2.0 g of chemical. You do not need to have **exactly** 2.00 g!!
- 2.) **Accurately** measure the mass of approximately 2.00 g of CaCl_2 into a 250 mL Erlenmeyer flask.
- 3.) **Accurately measure the mass of a filter paper.**
- 4.) Pour **approximately** 100.0 mL of distilled water into each beaker and swirl the contents until the chemical is completely dissolved. Carefully pour the CaCl_2 solution into the Na_2CO_3 solution and allow the reaction to "sit" for about 5 minutes. Filter the precipitate into the filter paper - make sure that you transfer all of the precipitate into the filter. Label the paper and set it aside until tomorrow.

Data and Observations -

- 1.) Record the mass of our Na_2CO_3 sample.
- 2.) Record the mass of our CaCl_2 sample.
- 3.) Record the mass of our filter paper